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Tri acidic bases are
those which have
three 'OH' group per

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unit as $\text{Al}(\text{OH})_3$ and $\text{Fe}(\text{OH})_3$, their one mole requires three mole of a base (NaOH) for complete neutralization. 365
366 367 Answer

Answers about Acids
and Bases

Liquids that measure 0-7 are acids with 0 being the strongest. Bases will measure

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from 7-14 with 14 being the strongest. A pH of 7 means that the liquid is neutral. Distilled water has a pH of 7. The scale is a convention for measurement and some sources argue that acids can be negative figures and bases can be stronger than 14.

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1. Both acids and
bases: A have a sour
taste B contain
hydrogen ions in
solution C are
corrosive 2. The pH

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of acids is: A greater
than 7 B equal to 7 C
less than 7 3. Bases
turn litmus paper: A
blue B purple C red 4.
The name given to a
soluble base is: A an
acid B a ...

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difference between an

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strength of an acid
using a titration. Skip
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Acid and Base
Worksheet - Answers.

1) Using your
knowledge of the
Brønsted-Lowry
theory of acids and
bases, write equations
for the following acid-
base reactions and

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indicate each
conjugate acid-base
pair: a) $\text{HNO}_3 + \text{OH}^-$
($\text{H}_2\text{O} + \text{NO}_3^-$).

HNO_3 and NO_3^-
make one pair OH^-
and H_2O make the
other. b) $\text{CH}_3\text{NH}_2 +$
 H_2O ($\text{CH}_3\text{NH}_3^+ +$
 OH^-

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Topic 12 - Acids,
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Level Chemistry

1 Acid = H_2O , base =
 NH_3 2 Acid = HCl ,
base = H_2O 3 Acid =
 HCOOH , base = KOH
4 Acid = HCl , base =
 CH_3COOH 5 Acid =
 HCl , base = NH_3 6
Acid = HCO_3^- , base
= OH^-

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Acids and bases

Acids, bases and

alkalis are found in

the laboratory and at

home. Acids and

bases can neutralise

each other. A base

that can dissolve in

water is also called an

alkali.

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2. Bases react with .
acids to produce salts
and water; water to
produce acids and
salts; salts to produce
acids and water;
neither acids, salts,
nor water

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multiple choice
questions and ...

an acid solution < 7 ; a
base solution > 7 ;

and; a neutral

solution = 7. (c) The
salts of strong acids
and weak bases give
acidic solution having
pH less than 7.

Example, NH_4Cl ,
Ammonium Chloride
will have pH less than

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7. The salts of weak acids and strong bases give basic solution having pH more than 7.

Acids, Bases and Salts
Class 10 Important
Questions with ...
Our general ideas
regarding acid and
alkali based on
Arrhenius concepts in
which an acid is a

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compound molecule that yield H^+ ion in water solution and the base is a compound molecule yields OH^- in water solution and utilization readily form solvent or water. Hence Arrhenius ' s theory very useful for explaining acid and bases in water solution but can ' t

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in glacial acetic acid
or ammonia solution.

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The chemistry of acids and bases and buffers is an important area. For example, the relative strengths of acids influences the formation of nitronium ions in the

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understanding of pH and buffers is essential in biology. Very early in the history of chemistry, many substances were designated as acids, bases, and salts.

Acids and Bases -
Definition, Examples,
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Acids Bases ;

Arrhenius Concept:

Acid is the substance when dissolved in water, increases the concentration of H^+ ions. The base is the substance when dissolved in water, increase the concentration of OH^- ions. Bronsted-Lowry Concept: Acids are the proton donor.

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acceptor. Lewis
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Chapter 19 Acids
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Key Acids And

Question: Describe
Bases Alphabet

the properties of acids and bases My answer is acids are substances which form H^+ (aq) ions when they are added to water. They are known as proton donors. Bases are also called alkalis, and are substances which form OH^- (aq) ions

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when added to water. They are known as proton acceptors. All acids have conjugate bases which are the particles that are left over when the acid ...

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