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Solubility product = K_{sp} Consider the equilibrium between undissolved silver chloride and a saturated silver chloride solution: $AgCl(s) \rightleftharpoons Ag^+(aq) + Cl^-(aq)$ Since this is a heterogeneous equilibrium, the equilibrium constant is $K_{sp} = [Ag^+][Cl^-]$ o always write the K_{sp} expression with the solid on the left

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At a certain temperature, the solubility of $Fe(OH)_2$ in water is 7.7×10^{-6} mol/L (M). Its K_{sp} can be calculated based on the equilibrium equation: $Fe(OH)_2 \rightleftharpoons Fe^{2+} + 2OH^-$ $Fe(OH)_2 \rightleftharpoons Fe^{2+} + 2OH^-$. Therefore, the solubility product expression is: $K_{sp} = [Fe^{2+}][OH^-]^2$ $K_{sp} = [Fe^{2+}][OH^-]^2$.

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